

Poly[tetraqua-di- μ_4 -malonato-barium(II)cadmium(II)]

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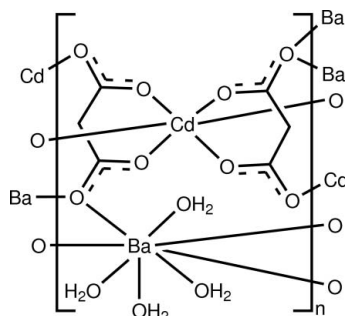
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Key indicators: single-crystal X-ray study; $T = 294$ K; mean $\sigma(\text{C}-\text{C}) = 0.019$ Å; R factor = 0.060; wR factor = 0.161; data-to-parameter ratio = 11.7.

In the title complex, $[\text{BaCd}(\text{C}_3\text{H}_2\text{O}_4)_2(\text{H}_2\text{O})_4]_n$, the Ba^{II} atoms, located on crystallographic twofold axes, adopt slightly distorted square-antiprismatic coordination geometries, while the Cd^{II} atoms, which lie on crystallographic centres of symmetry, have a distorted octahedral coordination. Each malonate dianion binds two different Cd^{II} atoms and two different Ba^{II} atoms. This connectivity generates alternating layers along $[100]$ in the structure, with one type containing Cd^{II} cations and malonate dianions, while the other is primarily composed of Ba^{II} ions and coordinated water molecules. The water molecules also participate in extensive $\text{O}-\text{H}\cdots\text{O}$ hydrogen bonding.

Related literature

For structural studies on the malonate dianion with its versatile coordination patterns, see: Delgado *et al.* (2004). For related structures, see Djeghri *et al.* (2005); Guo & Guo (2006).



Experimental

Crystal data

$[\text{BaCd}(\text{C}_3\text{H}_2\text{O}_4)_2(\text{H}_2\text{O})_4]$	$V = 1261.0$ (4) Å ³
$M_r = 525.90$	$Z = 4$
Orthorhombic, $Pccn$	Mo $K\alpha$ radiation
$a = 18.809$ (4) Å	$\mu = 4.85$ mm ⁻¹
$b = 6.9224$ (14) Å	$T = 294$ K
$c = 9.6849$ (19) Å	$0.24 \times 0.20 \times 0.10$ mm

Data collection

Bruker SMART CCD area-detector diffractometer	5716 measured reflections
Absorption correction: multi-scan (<i>SADABS</i> ; Sheldrick, 2000)	1103 independent reflections
$T_{\text{min}} = 0.370$, $T_{\text{max}} = 0.662$	978 reflections with $I > 2\sigma(I)$
	$R_{\text{int}} = 0.059$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.060$	94 parameters
$wR(F^2) = 0.161$	H-atom parameters constrained
$S = 1.06$	$\Delta\rho_{\text{max}} = 1.94$ e Å ⁻³
1103 reflections	$\Delta\rho_{\text{min}} = -1.24$ e Å ⁻³

Table 1

Selected bond lengths (Å).

Ba1—O4 ⁱ	2.794 (9)	Cd1—O2 ⁱⁱ	2.227 (10)
Ba1—O6	2.809 (10)	Cd1—O3	2.227 (9)
Ba1—O4	2.854 (9)	Cd1—O1 ⁱⁱⁱ	2.364 (8)
Ba1—O5	2.877 (10)		

Symmetry codes: (i) $x, -y + \frac{1}{2}, z + \frac{1}{2}$; (ii) $-x + 1, -y + 1, -z + 1$; (iii) $-x + 1, y + \frac{1}{2}, -z + \frac{1}{2}$.

Table 2

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
O6—H6B \cdots O5 ⁱⁱⁱ	0.87	2.08	2.893 (14)	157
O6—H6A \cdots O1 ^{iv}	0.85	1.99	2.781 (13)	156
O5—H5B \cdots O6 ⁱ	0.87	2.19	2.919 (15)	141
O5—H5A \cdots O2 ⁱⁱ	0.84	2.01	2.810 (14)	159

Symmetry codes: (i) $x, -y + \frac{1}{2}, z + \frac{1}{2}$; (ii) $-x + 1, -y + 1, -z + 1$; (iii) $x, y - 1, z$; (iv) $-x + 1, y - \frac{1}{2}, -z + \frac{1}{2}$.

Data collection: *SMART* (Bruker, 1997); cell refinement: *SAINTE* (Bruker, 1997); data reduction: *SAINTE*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: SJ5064).

References

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supplementary materials

Acta Cryst. (2011). E67, m23-m24 [doi:10.1107/S1600536810049780]

Poly[tetraaqua-di- μ_4 -malonato-barium(II)cadmium(II)]

M.-L. Guo, W.-J. Gao, C.-C. Luo and L. Liu

Comment

The malonate dianion, with two neighboring carboxylate groups, is a very flexible ligand. Its basic coordination mode is as a chelate *via* two distal carboxylate oxygen atoms to form a six-membered ring and the coordinating ability of the nonchelating oxygen atoms makes the formation of polymeric networks possible (Djeghri *et al.*, 2005; Guo & Guo, 2006). On the other hand, malonate can also coordinate in monodentate, chelated bidentate and bridging modes to create various molecular architectures (Delgado *et al.*, 2004). Herein, we report the structure of the title heterobimetallic malonate complex, (I). It and the chemically similar complex poly[tetraaqua-di- μ_4 -malonato-barium(II)zinc(II)] (Guo & Guo, 2006) are isotypic.

The asymmetric unit in the structure of (I) comprises half a Ba^{II} cation, half a Cd^{II} cation, a complete malonate dianion defined by C1—C3/O1—O4 and two independent water molecules involving O5 and O6. Fig. 1 shows a symmetry-expanded view which displays the full coordination of the Ba²⁺ and Cd²⁺ centers. Selected geometric parameters are given in Table 1.

The Ba²⁺ cation, lying on a crystallographic twofold axis, is eight-coordinate, bonded to oxygen atoms of four different malonate groups and four water molecules with Ba—O distances ranging from 2.793 (9) to 2.878 (10) Å. The Ba polyhedra may be described as slightly distorted square antiprisms. They share edges to form chains propagating along *c*.

The Cd²⁺ cations, lie on crystallographic centres of symmetry, and have distorted octahedral coordination, with O2 and O3 of two bidentate malonate anions at the equatorial sites and two O1 atoms from two other malonate anions at the apical sites.

Also evident in Fig. 1 is the variability of the coordination modes of the malonate dianion with monodentate (O1), bidentate chelating (O2 and O3) and bridging (O4) bonding modes all present.

The structure as a whole consists of two distinct types of layer, both parallel to (100) and stacked alternately in the direction of *a*. The first of these (Fig. 2) is composed entirely of Cd^{II} ions and malonate dianions and occurs at *x* = 0 and 1/2. The other type of layer, type 2, alternating with the first and centred on *x* = 1/4 and 3/4 contains, primarily, the Ba ions and the water molecules. Two forms of connectivity occur within the type 2 layers. First of all O4 atoms on the surfaces of the type 1 layers create chains of edge sharing Ba polyhedra propagating along *c* and at the same time link the two types of layer and complete the three-dimensional connectivity of the structure. The interlayer connectivity is further enhanced by the hydrogen bonds of the form O5—H5A \cdots O2^{iv} and O6—H6A \cdots O1^{vi} given in Table 2.

Experimental

The title complex was prepared under continuous stirring with successive addition of malonic acid (0.43 g, 4 mmol), cadmium(II) chloride (0.37 g, 2 mmol) and Ba(OH)₂·8H₂O (0.63 g, 2 mmol) to distilled water (40 ml) at room temperature. After filtration, slow evaporation over a period of a week at room temperature provided colorless plate-like crystals of (I).

Refinement

The H atoms of the water molecule were found in difference Fourier maps and during refinement were fixed at an O–H distance of 0.85 Å, and with $U_{\text{iso}}(\text{H}) = 1.2 U_{\text{eq}}(\text{O})$. The H atoms of C–H groups were placed geometrically and during refinement were treated using a riding model, with C–H = 0.97 Å, and with $U_{\text{iso}}(\text{H}) = 1.2 U_{\text{eq}}(\text{C})$.

Figures

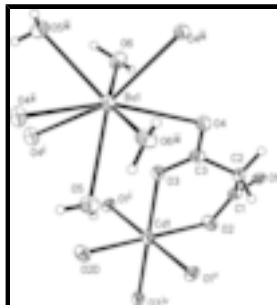


Fig. 1. The coordination of the metal ions in (I). Displacement ellipsoids are drawn at the 30% probability level. Symmetry codes (i) $x, -y + 1/2, z + 1/2$; (ii) $-x + 1/2, y, z + 1/2$; (iii) $-x + 1/2, -y + 1/2, z$; (iv) $-x + 1, -y + 1, -z + 1$; (v) $-x + 1, y + 1/2, -z + 1/2$.

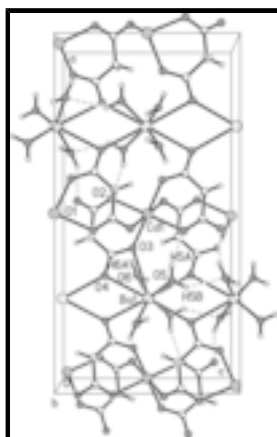


Fig. 2. A view, approximately along the b axis, showing the alternation of type 1 and type 2 layers along the a axis.

Poly[tetraaqua-di- μ_4 -malonato-barium(II)cadmium(II)]

Crystal data

[BaCd(C₃H₂O₄)₂(H₂O)₄]

$M_r = 525.90$

Orthorhombic, *Pccn*

Hall symbol: -P 2ab 2ac

$a = 18.809 (4) \text{ \AA}$

$b = 6.9224 (14) \text{ \AA}$

$c = 9.6849 (19) \text{ \AA}$

$V = 1261.0 (4) \text{ \AA}^3$

$Z = 4$

$F(000) = 992$

$D_x = 2.770 \text{ Mg m}^{-3}$

Mo $K\alpha$ radiation, $\lambda = 0.71073 \text{ \AA}$

Cell parameters from 3576 reflections

$\theta = 3.1\text{--}26.4^\circ$

$\mu = 4.85 \text{ mm}^{-1}$

$T = 294 \text{ K}$

Prism, colorless

$0.24 \times 0.20 \times 0.10 \text{ mm}$

Data collection

Bruker SMART CCD area-detector diffractometer	1103 independent reflections
Radiation source: sealed tube graphite	978 reflections with $I > 2\sigma(I)$
φ and ω scans	$R_{\text{int}} = 0.059$
Absorption correction: multi-scan (SADABS; Sheldrick, 2000)	$\theta_{\text{max}} = 25.0^\circ$, $\theta_{\text{min}} = 2.2^\circ$
$T_{\text{min}} = 0.370$, $T_{\text{max}} = 0.662$	$h = -22 \rightarrow 10$
5716 measured reflections	$k = -8 \rightarrow 8$
	$l = -10 \rightarrow 11$

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.060$	H-atom parameters constrained
$wR(F^2) = 0.161$	$w = 1/[\sigma^2(F_o^2) + (0.0364P)^2 + 63.5907P]$
$S = 1.06$	where $P = (F_o^2 + 2F_c^2)/3$
1103 reflections	$(\Delta/\sigma)_{\text{max}} = 0.001$
94 parameters	$\Delta\rho_{\text{max}} = 1.94 \text{ e } \text{\AA}^{-3}$
0 restraints	$\Delta\rho_{\text{min}} = -1.24 \text{ e } \text{\AA}^{-3}$
Primary atom site location: structure-invariant direct methods	Extinction correction: <i>SHELXL97</i> (Sheldrick, 2008), $F_c^* = kFc[1+0.001xFc^2\lambda^3/\sin(2\theta)]^{-1/4}$
	Extinction coefficient: 0.0147 (13)

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Ba1	0.2500	0.2500	0.50669 (11)	0.0320 (5)
Cd1	0.5000	0.5000	0.5000	0.0320 (6)
O1	0.5419 (5)	0.2940 (12)	0.0939 (9)	0.030 (2)
O2	0.5475 (5)	0.4394 (15)	0.2939 (10)	0.035 (2)

supplementary materials

O3	0.4014 (5)	0.3657 (15)	0.4162 (9)	0.034 (2)
O4	0.3196 (5)	0.3410 (17)	0.2530 (9)	0.035 (2)
O5	0.3078 (5)	0.5890 (15)	0.6370 (11)	0.041 (3)
H5B	0.2892	0.6078	0.7177	0.061*
H5A	0.3517	0.5665	0.6370	0.061*
O6	0.3163 (5)	-0.1005 (15)	0.4378 (11)	0.041 (2)
H6A	0.3571	-0.1227	0.4039	0.062*
H6B	0.3010	-0.1850	0.4966	0.062*
C1	0.5135 (7)	0.3835 (18)	0.1916 (14)	0.029 (3)
C2	0.4344 (7)	0.4323 (19)	0.1784 (13)	0.028 (3)
H2A	0.4307	0.5714	0.1686	0.033*
H2B	0.4178	0.3763	0.0924	0.033*
C3	0.3825 (7)	0.3720 (19)	0.2905 (15)	0.030 (3)

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Ba1	0.0311 (8)	0.0367 (8)	0.0281 (7)	0.0015 (5)	0.000	0.000
Cd1	0.0341 (9)	0.0357 (9)	0.0261 (8)	-0.0001 (6)	-0.0021 (6)	-0.0004 (6)
O1	0.035 (5)	0.018 (4)	0.037 (5)	-0.003 (4)	0.005 (4)	-0.005 (4)
O2	0.028 (5)	0.045 (6)	0.033 (5)	0.004 (5)	-0.006 (4)	-0.007 (5)
O3	0.031 (5)	0.046 (6)	0.026 (5)	-0.009 (4)	-0.002 (4)	0.003 (4)
O4	0.020 (5)	0.052 (6)	0.032 (5)	-0.010 (5)	-0.001 (4)	0.001 (4)
O5	0.030 (5)	0.041 (6)	0.052 (6)	0.001 (5)	-0.006 (5)	-0.001 (5)
O6	0.031 (5)	0.044 (6)	0.049 (6)	0.004 (5)	0.011 (5)	0.006 (5)
C1	0.031 (7)	0.021 (6)	0.034 (7)	0.001 (5)	0.000 (6)	0.001 (5)
C2	0.029 (7)	0.029 (7)	0.025 (6)	-0.003 (6)	0.001 (6)	0.000 (5)
C3	0.032 (7)	0.023 (7)	0.036 (7)	0.001 (6)	-0.002 (6)	-0.003 (6)

Geometric parameters (\AA , $^\circ$)

Ba1—O4 ⁱ	2.794 (9)	Cd1—O1 ^v	2.364 (8)
Ba1—O4 ⁱⁱ	2.794 (9)	Cd1—O1 ⁱ	2.364 (8)
Ba1—O6 ⁱⁱⁱ	2.809 (10)	O1—C1	1.252 (16)
Ba1—O6	2.809 (10)	O1—Cd1 ^{vi}	2.364 (8)
Ba1—O4 ⁱⁱⁱ	2.854 (9)	O2—C1	1.241 (16)
Ba1—O4	2.854 (9)	O3—C3	1.269 (17)
Ba1—O5	2.877 (10)	O4—C3	1.255 (16)
Ba1—O5 ⁱⁱⁱ	2.877 (10)	O4—Ba1 ^{vii}	2.794 (9)
Ba1—O3 ⁱⁱⁱ	3.086 (9)	O5—H5B	0.8658
Ba1—O3	3.086 (9)	O5—H5A	0.8410
Ba1—C3 ⁱⁱⁱ	3.363 (14)	O6—H6A	0.8479
Ba1—C3	3.363 (14)	O6—H6B	0.8659
Cd1—O2 ^{iv}	2.227 (10)	C1—C2	1.531 (18)
Cd1—O2	2.227 (10)	C2—C3	1.518 (19)
Cd1—O3 ^{iv}	2.227 (9)	C2—H2A	0.9700
Cd1—O3	2.227 (9)	C2—H2B	0.9700

O4 ⁱ —Ba1—O4 ⁱⁱ	62.7 (4)	O4 ⁱ —Ba1—C3	103.9 (3)
O4 ⁱ —Ba1—O6 ⁱⁱⁱ	127.3 (3)	O4 ⁱⁱ —Ba1—C3	145.4 (3)
O4 ⁱⁱ —Ba1—O6 ⁱⁱⁱ	78.5 (3)	O6 ⁱⁱⁱ —Ba1—C3	88.0 (3)
O4 ⁱ —Ba1—O6	78.5 (3)	O6—Ba1—C3	74.9 (3)
O4 ⁱⁱ —Ba1—O6	127.3 (3)	O4 ⁱⁱⁱ —Ba1—C3	81.9 (3)
O6 ⁱⁱⁱ —Ba1—O6	152.5 (4)	O4—Ba1—C3	21.3 (3)
O4 ⁱ —Ba1—O4 ⁱⁱⁱ	154.2 (5)	O5—Ba1—C3	77.8 (3)
O4 ⁱⁱ —Ba1—O4 ⁱⁱⁱ	124.7 (4)	O5 ⁱⁱⁱ —Ba1—C3	139.3 (3)
O6 ⁱⁱⁱ —Ba1—O4 ⁱⁱⁱ	77.4 (3)	O3 ⁱⁱⁱ —Ba1—C3	124.9 (3)
O6—Ba1—O4 ⁱⁱⁱ	79.0 (3)	O3—Ba1—C3	22.2 (3)
O4 ⁱ —Ba1—O4	124.7 (4)	C3 ⁱⁱⁱ —Ba1—C3	103.0 (5)
O4 ⁱⁱ —Ba1—O4	154.2 (5)	O2 ^{iv} —Cd1—O2	180.000 (1)
O6 ⁱⁱⁱ —Ba1—O4	79.0 (3)	O2 ^{iv} —Cd1—O3 ^{iv}	85.9 (3)
O6—Ba1—O4	77.4 (3)	O2—Cd1—O3 ^{iv}	94.1 (3)
O4 ⁱⁱⁱ —Ba1—O4	61.2 (4)	O2 ^{iv} —Cd1—O3	94.1 (3)
O4 ⁱ —Ba1—O5	68.4 (3)	O2—Cd1—O3	85.9 (3)
O4 ⁱⁱ —Ba1—O5	67.6 (3)	O3 ^{iv} —Cd1—O3	180.0 (3)
O6 ⁱⁱⁱ —Ba1—O5	64.4 (3)	O2 ^{iv} —Cd1—O1 ^v	92.8 (3)
O6—Ba1—O5	129.8 (3)	O2—Cd1—O1 ^v	87.2 (3)
O4 ⁱⁱⁱ —Ba1—O5	136.9 (3)	O3 ^{iv} —Cd1—O1 ^v	93.3 (3)
O4—Ba1—O5	91.4 (3)	O3—Cd1—O1 ^v	86.7 (3)
O4 ⁱ —Ba1—O5 ⁱⁱⁱ	67.6 (3)	O2 ^{iv} —Cd1—O1 ⁱ	87.2 (3)
O4 ⁱⁱ —Ba1—O5 ⁱⁱⁱ	68.4 (3)	O2—Cd1—O1 ⁱ	92.8 (3)
O6 ⁱⁱⁱ —Ba1—O5 ⁱⁱⁱ	129.8 (3)	O3 ^{iv} —Cd1—O1 ⁱ	86.7 (3)
O6—Ba1—O5 ⁱⁱⁱ	64.4 (3)	O3—Cd1—O1 ⁱ	93.3 (3)
O4 ⁱⁱⁱ —Ba1—O5 ⁱⁱⁱ	91.4 (3)	O1 ^v —Cd1—O1 ⁱ	180.0 (4)
O4—Ba1—O5 ⁱⁱⁱ	136.9 (3)	C1—O1—Cd1 ^{vi}	125.1 (8)
O5—Ba1—O5 ⁱⁱⁱ	128.0 (4)	C1—O2—Cd1	124.6 (9)
O4 ⁱ —Ba1—O3 ⁱⁱⁱ	128.1 (3)	C3—O3—Cd1	124.7 (9)
O4 ⁱⁱ —Ba1—O3 ⁱⁱⁱ	82.4 (3)	C3—O3—Ba1	91.3 (8)
O6 ⁱⁱⁱ —Ba1—O3 ⁱⁱⁱ	75.3 (3)	Cd1—O3—Ba1	140.6 (4)
O6—Ba1—O3 ⁱⁱⁱ	96.8 (3)	C3—O4—Ba1 ^{vii}	136.6 (9)
O4 ⁱⁱⁱ —Ba1—O3 ⁱⁱⁱ	43.5 (3)	C3—O4—Ba1	102.8 (8)
O4—Ba1—O3 ⁱⁱⁱ	103.7 (3)	Ba1 ^{vii} —O4—Ba1	118.0 (3)
O5—Ba1—O3 ⁱⁱⁱ	133.2 (3)	Ba1—O5—H5B	111.5
O5 ⁱⁱⁱ —Ba1—O3 ⁱⁱⁱ	64.2 (3)	Ba1—O5—H5A	102.7
O4 ⁱ —Ba1—O3	82.4 (3)	H5B—O5—H5A	115.1
O4 ⁱⁱ —Ba1—O3	128.1 (3)	Ba1—O6—H6A	130.6
O6 ⁱⁱⁱ —Ba1—O3	96.8 (3)	Ba1—O6—H6B	106.2
O6—Ba1—O3	75.3 (3)	H6A—O6—H6B	115.7
O4 ⁱⁱⁱ —Ba1—O3	103.7 (3)	O2—C1—O1	122.6 (12)

supplementary materials

O4—Ba1—O3	43.5 (3)	O2—C1—C2	119.9 (12)
O5—Ba1—O3	64.2 (3)	O1—C1—C2	117.4 (12)
O5 ⁱⁱⁱ —Ba1—O3	133.2 (3)	C3—C2—C1	120.2 (11)
O3 ⁱⁱⁱ —Ba1—O3	147.0 (3)	C3—C2—H2A	107.3
O4 ⁱ —Ba1—C3 ⁱⁱⁱ	145.3 (3)	C1—C2—H2A	107.3
O4 ⁱⁱ —Ba1—C3 ⁱⁱⁱ	103.9 (3)	C3—C2—H2B	107.3
O6 ⁱⁱⁱ —Ba1—C3 ⁱⁱⁱ	74.9 (3)	C1—C2—H2B	107.3
O6—Ba1—C3 ⁱⁱⁱ	88.0 (3)	H2A—C2—H2B	106.9
O4 ⁱⁱⁱ —Ba1—C3 ⁱⁱⁱ	21.3 (3)	O4—C3—O3	122.4 (13)
O4—Ba1—C3 ⁱⁱⁱ	81.9 (3)	O4—C3—C2	116.4 (12)
O5—Ba1—C3 ⁱⁱⁱ	139.3 (3)	O3—C3—C2	121.0 (12)
O5 ⁱⁱⁱ —Ba1—C3 ⁱⁱⁱ	77.8 (3)	O4—C3—Ba1	55.9 (7)
O3 ⁱⁱⁱ —Ba1—C3 ⁱⁱⁱ	22.2 (3)	O3—C3—Ba1	66.6 (7)
O3—Ba1—C3 ⁱⁱⁱ	124.9 (3)	C2—C3—Ba1	172.1 (9)
O3 ^{iv} —Cd1—O2—C1	-170.4 (11)	O6—Ba1—O4—Ba1 ^{vii}	84.2 (4)
O3—Cd1—O2—C1	9.6 (11)	O4 ⁱⁱⁱ —Ba1—O4—Ba1 ^{vii}	0.0
O1 ^v —Cd1—O2—C1	-77.2 (11)	O5—Ba1—O4—Ba1 ^{vii}	-145.1 (4)
O1 ⁱ —Cd1—O2—C1	102.8 (11)	O5 ⁱⁱⁱ —Ba1—O4—Ba1 ^{vii}	56.8 (6)
O2 ^{iv} —Cd1—O3—C3	148.6 (11)	O3 ⁱⁱⁱ —Ba1—O4—Ba1 ^{vii}	-9.8 (5)
O2—Cd1—O3—C3	-31.4 (11)	O3—Ba1—O4—Ba1 ^{vii}	166.0 (7)
O1 ^v —Cd1—O3—C3	56.0 (11)	C3 ⁱⁱⁱ —Ba1—O4—Ba1 ^{vii}	-5.5 (4)
O1 ⁱ —Cd1—O3—C3	-124.0 (11)	C3—Ba1—O4—Ba1 ^{vii}	164.9 (12)
O2 ^{iv} —Cd1—O3—Ba1	-3.7 (7)	Cd1—O2—C1—O1	-156.7 (9)
O2—Cd1—O3—Ba1	176.3 (7)	Cd1—O2—C1—C2	26.4 (17)
O1 ^v —Cd1—O3—Ba1	-96.3 (6)	Cd1 ^{vi} —O1—C1—O2	129.4 (11)
O1 ⁱ —Cd1—O3—Ba1	83.7 (6)	Cd1 ^{vi} —O1—C1—C2	-53.6 (14)
O4 ⁱ —Ba1—O3—C3	166.0 (8)	O2—C1—C2—C3	-57.2 (17)
O4 ⁱⁱ —Ba1—O3—C3	-147.8 (8)	O1—C1—C2—C3	125.7 (13)
O6 ⁱⁱⁱ —Ba1—O3—C3	-67.2 (8)	Ba1 ^{vii} —O4—C3—O3	-162.7 (10)
O6—Ba1—O3—C3	86.0 (8)	Ba1—O4—C3—O3	-2.2 (15)
O4 ⁱⁱⁱ —Ba1—O3—C3	11.5 (8)	Ba1 ^{vii} —O4—C3—C2	22 (2)
O4—Ba1—O3—C3	-1.1 (7)	Ba1—O4—C3—C2	-177.9 (9)
O5—Ba1—O3—C3	-124.4 (9)	Ba1 ^{vii} —O4—C3—Ba1	-160.5 (15)
O5 ⁱⁱⁱ —Ba1—O3—C3	116.7 (8)	Cd1—O3—C3—O4	-160.9 (10)
O3 ⁱⁱⁱ —Ba1—O3—C3	6.3 (7)	Ba1—O3—C3—O4	2.0 (14)
C3 ⁱⁱⁱ —Ba1—O3—C3	9.2 (11)	Cd1—O3—C3—C2	14.7 (18)
O4 ⁱ —Ba1—O3—Cd1	-36.5 (7)	Ba1—O3—C3—C2	177.5 (11)
O4 ⁱⁱ —Ba1—O3—Cd1	9.7 (8)	Cd1—O3—C3—Ba1	-162.9 (10)
O6 ⁱⁱⁱ —Ba1—O3—Cd1	90.4 (7)	C1—C2—C3—O4	-151.1 (13)
O6—Ba1—O3—Cd1	-116.5 (7)	C1—C2—C3—O3	33.1 (19)
O4 ⁱⁱⁱ —Ba1—O3—Cd1	169.0 (6)	O4 ⁱ —Ba1—C3—O4	167.7 (7)

O4—Ba1—O3—Cd1	156.5 (9)	O4 ⁱⁱ —Ba1—C3—O4	-130.5 (10)
O5—Ba1—O3—Cd1	33.2 (6)	O6 ⁱⁱⁱ —Ba1—C3—O4	-64.3 (9)
O5 ⁱⁱⁱ —Ba1—O3—Cd1	-85.7 (7)	O6—Ba1—C3—O4	94.0 (9)
O3 ⁱⁱⁱ —Ba1—O3—Cd1	163.8 (7)	O4 ⁱⁱⁱ —Ba1—C3—O4	13.3 (10)
C3 ⁱⁱⁱ —Ba1—O3—Cd1	166.7 (6)	O5—Ba1—C3—O4	-128.5 (9)
C3—Ba1—O3—Cd1	157.5 (13)	O5 ⁱⁱⁱ —Ba1—C3—O4	96.0 (9)
O4 ⁱ —Ba1—O4—C3	-14.5 (8)	O3 ⁱⁱⁱ —Ba1—C3—O4	6.2 (10)
O4 ⁱⁱ —Ba1—O4—C3	84.2 (9)	O3—Ba1—C3—O4	-178.0 (14)
O6 ⁱⁱⁱ —Ba1—O4—C3	113.5 (9)	C3 ⁱⁱⁱ —Ba1—C3—O4	9.7 (8)
O6—Ba1—O4—C3	-80.7 (9)	O4 ⁱ —Ba1—C3—O3	-14.3 (8)
O4 ⁱⁱⁱ —Ba1—O4—C3	-164.9 (12)	O4 ⁱⁱ —Ba1—C3—O3	47.5 (10)
O5—Ba1—O4—C3	49.9 (9)	O6 ⁱⁱⁱ —Ba1—C3—O3	113.7 (8)
O5 ⁱⁱⁱ —Ba1—O4—C3	-108.1 (9)	O6—Ba1—C3—O3	-88.0 (8)
O3 ⁱⁱⁱ —Ba1—O4—C3	-174.8 (9)	O4 ⁱⁱⁱ —Ba1—C3—O3	-168.7 (8)
O3—Ba1—O4—C3	1.1 (8)	O4—Ba1—C3—O3	178.0 (14)
C3 ⁱⁱⁱ —Ba1—O4—C3	-170.4 (8)	O5—Ba1—C3—O3	49.5 (8)
O4 ⁱ —Ba1—O4—Ba1 ^{vii}	150.4 (5)	O5 ⁱⁱⁱ —Ba1—C3—O3	-86.0 (9)
O4 ⁱⁱ —Ba1—O4—Ba1 ^{vii}	-110.9 (6)	O3 ⁱⁱⁱ —Ba1—C3—O3	-175.8 (5)
O6 ⁱⁱⁱ —Ba1—O4—Ba1 ^{vii}	-81.6 (4)	C3 ⁱⁱⁱ —Ba1—C3—O3	-172.3 (9)

Symmetry codes: (i) $x, -y+1/2, z+1/2$; (ii) $-x+1/2, y, z+1/2$; (iii) $-x+1/2, -y+1/2, z$; (iv) $-x+1, -y+1, -z+1$; (v) $-x+1, y+1/2, -z+1/2$; (vi) $-x+1, y-1/2, -z+1/2$; (vii) $-x+1/2, y, z-1/2$.

Hydrogen-bond geometry (Å, °)

<i>D</i> —H \cdots <i>A</i>	<i>D</i> —H	H \cdots <i>A</i>	<i>D</i> \cdots <i>A</i>	<i>D</i> —H \cdots <i>A</i>
O6—H6B \cdots O5 ^{viii}	0.87	2.08	2.893 (14)	157
O6—H6A \cdots O1 ^{vi}	0.85	1.99	2.781 (13)	156
O5—H5B \cdots O6 ⁱ	0.87	2.19	2.919 (15)	141
O5—H5A \cdots O2 ^{iv}	0.84	2.01	2.810 (14)	159

Symmetry codes: (viii) $x, y-1, z$; (vi) $-x+1, y-1/2, -z+1/2$; (i) $x, -y+1/2, z+1/2$; (iv) $-x+1, -y+1, -z+1$.

Fig. 1

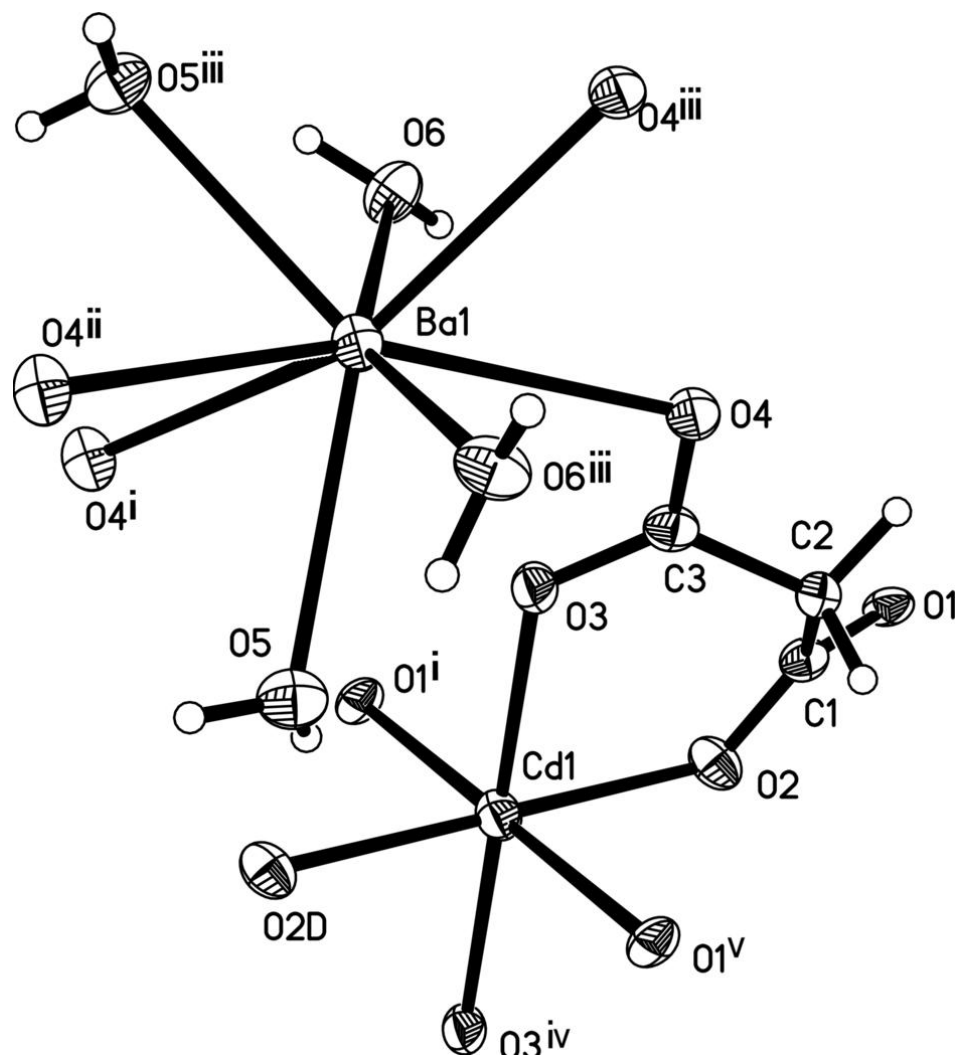


Fig. 2

